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Section 15-5

Spontaneous Processes

A

spontaneous process is

a physical or chemical

change that once

begun, occurs with no

outside intervention. </

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Many spontaneous processes require some energy from the surroundings to start the process.

53.

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Terms in this set (75)
boiling point elevation.

The temperature
difference between a
solution's boiling point
and a pure solvent's

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Figure 15.4.1 Energy
Changes

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Thermite Reaction
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state function, the
overall enthalpy
change for the reaction
of 2 mol of Al(s) with 1
mol of Fe₂O₃(s) is
−851.1 kJ, whether the
reaction occurs in a
single step (ΔH 4,
shown on the left) or in
three hypothetical
steps (shown on the
right) that involve the
successive formation of
solid Al₂O₃ and ...

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experts! The volume of
Oxygen and nitrogen
gas required to
produce 15.4 L
nitrogen dioxide needs
to be determined.

**The volume of
Oxygen and nitrogen
gas required to
produce ...**

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your textbook, read about the nature of energy In the space at the left, write true if the statement is true; if the statement is false, change ...

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